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## Chapter2 Acids Bases And Salts

Acid + base  $\rightarrow$  salt + water + heat.  $H_2SO_4 + MgO \rightarrow MgSO_4 + H_2O$   
 $2HCl + Mg(OH)_2 \rightarrow MgCl_2 + 2H_2O$ . 2. Reaction of non-metal oxides with bases. Non-metal oxides are acidic in nature  
Base + Non-metal oxide  $\rightarrow$  salt + water + heat.  
 $2NaOH + CO_2 \rightarrow Na_2CO_3 + H_2O$ . To know more about Properties of Acids and Bases, visit here. Water Acids and bases in water

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Class 10 Chapter 2 Science Notes Acids, Bases, and Salts  
The reaction between an acid and a base to give a salt and water is known as a neutralisation reaction. In general, a neutralisation reaction can be written as  $\text{Base} + \text{Acid} \rightarrow \text{Salt} + \text{Water}$   
2.1.5 Reaction of Metallic Oxides with Acids Activity  
2.7 n Take a small amount of copper oxide in a beaker and add dilute hydrochloric acid slowly while stirring.

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## CHAPTER2 Acids, Bases and Salts

CBSE Class 10 Science Notes Chapter 2 Acids Bases and Salts. 1. Common Salt (Sodium Chloride): Sodium chloride (NaCl) is also known as Common or Table Salt. It is formed after the reaction between sodium ... 2. Bleaching Powder (CaOCl<sub>2</sub>): Bleaching powder is also known as chloride of lime. It is a ...

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Acids Bases and Salts Class 10 Notes Science Chapter 2 ...  
Notes Chapter - 2 Acids, Bases and Salts  
ACIDS: □ These are the substances which have sour in taste. □ They turns blue litmus solution to red. □ They give H ions in aqueous solution  
Examples of Acids : - HCl - Hydrochloric Acid  
BASES □ These substances are bitter in taste. □ They turns red litmus solution to blue.

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## Chapter - 2, Acids Bases and Salts Notes

CBSE Class 10 Science Chapter 2 Notes Acids Bases and Salts  
Acids Bases and Salts Class 10 Notes Understanding the Lesson 1. Acids and Bases: Acids are sour in taste and change the colour of blue litmus to red.

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Acids Bases and Salts Class 10 Notes Science Chapter 2  
Salts . When acid and base neutralize, salts are formed.  
Strong acid and strong base combines to form neutral salt.  
NaOH + HCl □ NaCl + H<sub>2</sub>O. Eq.1. Formation of Neutral Salt.  
Strong acid and weak base combine to form acidic salt. For Example, Hydrochloric Acid and ammonium hydroxide

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combine to form ammonium chloride. Other examples, sodium hydrogen carbonate, sodium hydrogen sulphate etc.

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Revision Notes for Science Chapter 2 - Acids, bases and ...  
Study Material and Notes of Ch 2 Acids, Bases and Salts  
Class 10th Science. □ The sour and bitter tastes of food are due to acids and bases are present in them. □ Acids are sour in taste and change the colour of blue litmus to red. □ Litmus solution is a purple dye, which is extracted from lichen.

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Notes of Ch 2 Acids, Bases and Salts | Class 10th Science  
Chemicals from common salt; NCERT Solutions for Class 10  
Chapter 2 Science Acids bases and salts. In this chapter, students get hold of basic knowledge on Acids bases and salts. In this chapter, various chemical properties of acids, bases and salts and the reaction of them with metals, non-metals are studied.

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NCERT Solutions for class 10 chapter 2 Acids, bases and salt  
When an acid and base react with each other to give a salt and water, it is termed as neutralisation reaction i.e., Acid + Base □ Salt + Water For example:  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ .  $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$ . Question 15. Give two important uses of washing soda and baking soda. Answer: Important uses of washing soda are:

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(ii) Acids react with metal carbonate and metals hydrogen carbonate to give salt, water and carbon dioxide. (iii) Acids react with bases to give salt and water. This reaction is called neutralization reaction. (iv) Acids react with metals oxides to give salt and water.

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Class 10 Chemistry Chapter 2 Important Questions with Answers Acids, Bases and Salts Acids, Bases and Salts

Class 10 Important Questions Very Short Answer Type

Question 1. How is the concentration of hydronium ( $\text{H}_3\text{O}^+$ ) ions affected when a solution of an acid is diluted? (2017 OD)

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Acids, Bases and Salts Class 10 Important Questions with ...

Video Timeline: 0:00:00 start 0:00:21 Difference between

Acids & bases 0:10:38 what are indicators? 0:17:07 Universal

Indicators 0:21:11 Salts 0:24:51 a) Sod...

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Salts dissociate in water fully and form their basic and acidic  
counterpart. For example here  $\text{CH}_3\text{COONa}$  will form  
 $\text{CH}_3\text{COOH}$  and  $\text{NaOH}$  with water. Since  $\text{NaOH}$  is highly  
basic, the solution will be basic. N.b. the reverse of this also  
applies. For example, a solution of a salt of a strong acid and  
weak base will be acidic.

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Acids Bases and Salts MCQ Class 10 Science chapter 2 -  
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Class 10 Science Chapter 2 Extra Questions and Answers  
Acids Bases and Salts. Extra Questions for Class 10 Science  
Chapter 2 Acids Bases and Salts with Answers Solutions.  
Acids Bases and Salts Extra Questions Very Short Answer  
Type. Question 1. Two solutions A and B have pH values of 5  
and 8 respectively. Which solution will be basic in nature ...

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Acids Bases and Salts Class 10 Extra Questions with ...  
For example, Hydrochloric acid ( $\text{HCl}$ ), Sulphuric acid ( $\text{H}_2\text{SO}_4$ ),  
Nitric acid ( $\text{HNO}_3$ ). Bases - Bases are compounds which  
give hydroxide ion in water solution. For example, Sodium  
hydroxide ( $\text{NaOH}$ ), Potassium hydroxide ( $\text{KOH}$ ), Calcium  
hydroxide ( $\text{Ca(OH)}_2$ ). Salts - Salts are compounds obtained

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by replacement of H from acid by a metal.

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