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1 point for answer) in 100.0 g sample 1.00 g / mol 85.7 g 12.0 g/ mol 14.3 mol H 7.14 mol C - 6.0 6 (CH2) 14.3 85.7% c 84.0 u = 14.3 mol H = 7.14 mol C 2 = CH2 emperical formula 6 12 (1 point) (7 point for method, 1 point for answer for each part) 12+2 u 10. molecules x - - 8 molecules H2O - 9.0 mol CO2 3.0 mol - -x + 18 g. 11 g +40. g -

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